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Electron deficiency - Wikipedia

Electron deficiency occurs when a compound has too few valence electrons for the connections between atoms to be described as covalent bonds. Electron deficient bonds are often better described as 3-center-2-electron bonds.

Electron-deficient compound | Britannica.com

In chemical bonding: Electron-deficient compounds

Another type of exception to the Lewis approach to bonding is the existence of compounds that possess too few electrons for a Lewis structure to be written.

Why is BCl₃ an electron-deficient compound, and why does ...

(AP or SAT II Chemistry level) Chlorine (3.16) is more electronegative than boron (2.04) so boron becomes relatively electron-deficient and finally have a room for foreign lone pair electron, thus forming a chemical bond. **Electron deficient molecules: three-center, two-electron ...**

Molecules containing less than two valence electrons per bond are referred to as electron deficient. The simplest such molecule is the trigonal trihydrogen cation (H₃⁺).

The bonding in this molecule is interpreted in terms of a symmetric three-center (3c) bonding molecular orbital containing the two valence electrons.

Electron Deficient Compounds (eBook, 1971)

[WorldCat.org]

This book is about compounds such as the boron hydrides and associated metal hydrides and alkyls which acquired the label 'electron deficient' when they were thought to contain too few valence Read more

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This book is about compounds such as the boron hydrides and associated metal hydrides and alkyls which acquired the label 'electron deficient' when they were thought to contain too few valence electrons to hold together.

Why is BF₃ an electron deficient compound? - Quora

The 1 unpaired electron find its pair with 1 electron of B each. As B have 3 electrons, 3 F find their pairs with each electron of B. So B now have 6 electrons in its vicinity (3 electron pairs) but it needs 4 pairs i.e. 8 electrons. So it is electron deficient.

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HALOALKANES (HALOGENOALKANES) are attracted to the slightly positive (electron deficient) carbon. Basic mechanism: the nucleophile uses its lone pair to provide the electrons for a new bond.

Other examples of electron deficient bonding |

SpringerLink

Abstract. The criterion of electron deficiency that has been used in this book is that a molecule or ion possesses too few bonding electrons to allow a pair to be allocated to every pair of atoms close enough to be regarded as covalently bonded.